

**(2E)-2-(2,4-Dichlorophenylsulfonyl)-3-(3-methoxyanilino)-3-(methylsulfonyl)acrylonitrile**Mario V. Capparelli,<sup>a\*</sup> Arthur R. Barazarte<sup>b</sup> and Jaime E. Charris<sup>b</sup><sup>a</sup>Escuela de Química, Facultad de Ciencias, Universidad Central de Venezuela, Caracas 1051, Venezuela, and <sup>b</sup>Facultad de Farmacia, Universidad Central de Venezuela, Caracas 1051, Venezuela

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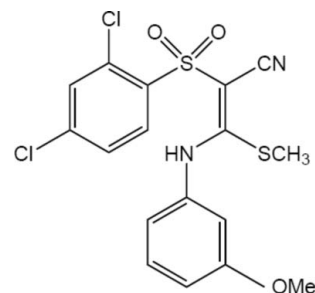
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Key indicators: single-crystal X-ray study;  $T = 296$  K; mean  $\sigma(\text{C}-\text{C}) = 0.003$  Å;  $R$  factor = 0.044;  $wR$  factor = 0.120; data-to-parameter ratio = 19.7.

The title compound,  $\text{C}_{17}\text{H}_{14}\text{Cl}_2\text{N}_2\text{O}_3\text{S}_2$ , and the 4-methoxyanilino analogue reported in the following paper have been used as starting materials to develop benzothiazine derivatives with antimalarial activity. The molecule displays an *E* (*trans*) configuration about the central double bond. Due to conjugation in the  $\text{C}=\text{C}-\text{C}\equiv\text{N}$  group, the putative single bond shows a significant shortening [1.421 (3) Å]. The molecule has a six-membered ring involving an intramolecular  $\text{N}-\text{H}\cdots\text{O}(\text{sulfonyl})$  bond, which is an example of resonance-assisted hydrogen bonding. There is also an intramolecular  $\text{N}-\text{H}\cdots\text{Cl}$  hydrogen bond. In the crystal structure, bonds of the  $\text{C}-\text{H}\cdots\text{O}(\text{sulfonyl})$  type form chains that run along [101], while  $\text{N}-\text{H}\cdots\text{O}(\text{sulfonyl})$  bonds connect centrosymmetrically related molecules in pairs of these chains, forming ribbons. Comparison of the  $\text{N}\cdots\text{O}$  distances in the intra- and intermolecular  $\text{N}-\text{H}\cdots\text{O}(\text{sulfonyl})$  bonds reveals that the  $\pi$ -bond co-operativity results in a strengthening of the intramolecular hydrogen bond. There are also  $\pi-\pi$  interactions between benzene rings of pairs of centrosymmetrically related molecules [centroid-centroid distance = 3.8612 (13) Å], as well as  $\text{C}-\text{H}\cdots\pi$  interactions.

**Related literature**

For related literature, see: Allen (2002); Allen *et al.* (1987); Barazarte *et al.* (2008); Capparelli *et al.* (2008); Charris *et al.* (2005, 2007); Gilli *et al.* (1989); Kennard *et al.* (2003); Krivokolysko *et al.* (2002); Song *et al.* (2005); Tominaga *et al.* (1989, 2002).

**Experimental***Crystal data*

$\text{C}_{17}\text{H}_{14}\text{Cl}_2\text{N}_2\text{O}_3\text{S}_2$   
 $M_r = 429.32$   
 Monoclinic,  $P2_1/n$   
 $a = 11.6125$  (7) Å  
 $b = 10.1419$  (6) Å  
 $c = 16.4048$  (10) Å  
 $\beta = 100.926$  (1)°

$V = 1897.0$  (2) Å<sup>3</sup>  
 $Z = 4$   
 Mo  $K\alpha$  radiation  
 $\mu = 0.58$  mm<sup>-1</sup>  
 $T = 296$  (2) K  
 $0.29 \times 0.22 \times 0.17$  mm

*Data collection*

Bruker SMART APEX  
 diffractometer  
 Absorption correction: multi-scan  
 (SADABS; Bruker, 2001)  
 $T_{\min} = 0.796$ ,  $T_{\max} = 0.906$

12828 measured reflections  
 4659 independent reflections  
 3653 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.024$

*Refinement*

$R[F^2 > 2\sigma(F^2)] = 0.044$   
 $wR(F^2) = 0.119$   
 $S = 1.04$   
 4659 reflections

237 parameters  
 H-atom parameters constrained  
 $\Delta\rho_{\max} = 0.58$  e Å<sup>-3</sup>  
 $\Delta\rho_{\min} = -0.27$  e Å<sup>-3</sup>

**Table 1**

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{N2}-\text{H2}\cdots\text{O2}$	0.86	2.14	2.756 (2)	129
$\text{N2}-\text{H2}\cdots\text{O2}^i$	0.86	2.31	3.005 (2)	138
$\text{N2}-\text{H2}\cdots\text{Cl1}$	0.86	2.72	3.2763 (18)	123
$\text{C4}-\text{H4C}\cdots\text{O1}^{\text{ii}}$	0.96	2.46	3.215 (3)	135
$\text{C4}-\text{H4B}\cdots\text{Cg2}$	0.96	2.74	3.501 (3)	137

Symmetry codes: (i)  $-x + 2, -y + 1, -z$ ; (ii)  $x + \frac{1}{2}, -y + \frac{1}{2}, z + \frac{1}{2}$ . Cg2 is the centroid of the C21-C26 ring.

Data collection: SMART (Bruker, 2002); cell refinement: SAINT (Bruker, 2001); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: ORTEP-3 (Farrugia, 1997); software used to prepare material for publication: WinGX (Farrugia, 1999) and PLATON (Spek, 2003).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: BG2184).

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**supplementary materials**

*Acta Cryst.* (2008). E64, o981-o982 [ doi:10.1107/S160053680801252X ]

**(2E)-2-(2,4-Dichlorophenylsulfonyl)-3-(3-methoxyanilino)-3-(methylsulfonyl)acrylonitrile**

**M. V. Capparelli, A. R. Barazarte and J. E. Charris**

**Comment**

The exploration of simple molecules with different functionalities for the synthesis of heterocycles is a worthwhile contribution to the chemistry of these compounds. The title compound (I), the 4"-methyl analogue (II) [see next paper: Capparelli *et al.*, 2008], and similar derivatives, have been used as effective synthons in the syntheses of some 1H-pyrrole-2,5-diones (Tominaga *et al.*, 2002), pyrimidine derivatives (Tominaga *et al.*, 1989) and 4H-1,4-benzothiazine-1,1-dioxides (Charris *et al.*, 2005). We used (I) and (II) as starting materials to develop benzothiazine derivatives with antimalarial activity (Charris *et al.*, 2007; Barazarte *et al.*, 2008).

The X-ray structure determination showed that there is one molecule per asymmetric unit (Fig. 1), which displays E (*trans*) configuration about the C2=C3 double bond. A search of the Cambridge Structural Database (version 5.29, updated Jan 2008) (Allen, 2002) produced no structures with the same central fragment (*i.e.* excluding the phenyl rings) of (I) for proper comparison, but a search for the more restricted fragment  $X-C(CN)=C(SMe)-N(H)-Y$  gave three comparable structures, *viz.* TAKDOZ (Krivokolysko *et al.*, 2002), AJULUM (Kennard *et al.*, 2003) and DALVES (Song *et al.*, 2005). Due to conjugation in the C3=C2—C1≡N1 moiety, the putative single bond C2—C1 shows a significant shortening, similar to the range 1.415 (7)–1.437 (4) Å observed in the aforementioned structures. Bond lengths (see Supplementary Materials) are in good agreement with the tabulated values (Allen *et al.*, 1987), except C24—C25 (1.359 (4) Å), which is *ca* 0.02 Å shorter than expected.

The title compound displays a six-membered ring involving an intramolecular N—H⋯O(sulfonyl) bond (Table 1), which is an example of resonance-assisted hydrogen bonding (RAHB) (Gilli *et al.*, 1989), as suggested by the ring bond lengths. Comparison with AJULUM and DALVES, which display similar rings [with N—H⋯O(carbonyl) and N—H⋯O(carboxyl) bonds respectively] and TAKDOZ, which does not have RAHB, reveal lengthenings of the C=C distances [AJULUM, 1.371 (2) Å; DALVES, 1.386 (7) Å; TAKDOZ, 1.345 (4) Å] and shortenings of the C—N bonds [AJULUM, 1.360 (2) Å; DALVES, 1.333 (6) Å; TAKDOZ, 1.404 (4) Å]. The situation is less clear in the S=O bond, with S1=O2 (involved in RAHB) just significantly longer than S1=O1. In addition to the RAHB ring, there is an N—H⋯Cl intramolecular bond. The molecular geometry also allows for a possible C4—H4b⋯Cg2 intramolecular interaction (Cg2: see below).

In the crystal structure (Fig. 2) bonds of the C—H⋯O(sulfonyl) type form chains that run along [1 0 1], while N—H⋯O(sulfonyl) bonds connect centrosymmetrically related atoms of pairs of these chains, to form ribbons. Comparison of the N⋯O distances of the intra- and intermolecular N—H⋯O(sulfonyl) bonds reveals that the  $\pi$ -bond cooperativity results in a strengthening of the intramolecular hydrogen bond, as indicated by the significant shortening of the corresponding N⋯O distance. There are also  $\pi$ - $\pi$  interactions between phenyl rings of pairs of centrosymmetrically related molecules, with Cg1⋯Cg2(2 - x, 1 - y, -z), 3.8612 (13) Å (Cgm: centroid of ring Cm1—Cm6, m = 1, 2).

## Experimental

To a solution of 2,4-dichlorobenzenesulfonylacetonitrile (1.50 g, 6.1 mmol) and KOH (0.46 g, 8.1 mmol) in anhydrous dioxane (10 ml) at 0 °C, was added dropwise 3-methoxyphenylisothiocyanate (1.00 g, 6.1 mmol) dissolved in anhydrous dioxane (3 ml). The resulting solution was stirred for 4 h at room temperature and then was added iodomethane (0.87 g, 6.1 mmol) and the mixture was stirred by 5 h at room temperature. The solvent was removed under reduced pressure and the residue was dissolved in dichloromethane (20 ml), washed with water, brine, dried over anhydrous sodium sulfate, filtered and concentrated to dryness. The solid was purified by recrystallization from ethanol. Yield: 1.91 g (74%). Crystals suitable for X-ray analysis were obtained during the recrystallization.

The IR spectrum was determined on a Shimadzu model 470 spectrophotometer. IR data [KBr pellets,  $\nu$  ( $\text{cm}^{-1}$ ): 3264 (NH), 2192 (CN), 1341 ( $\text{SO}_2$ ), 1133 ( $\text{SO}_2$ ).

The  $^1\text{H}$  NMR spectrum was recorded using a Jeol Eclipse 270 MHz [ $\text{CDCl}_3/\text{TMS}$ ,  $\delta$  (p.p.m.), atomic numbering as in Fig. 1]: 2.21 (s, 3H,  $\text{SCH}_3$ ), 3.81 (s, 3H,  $\text{OCH}_3$ ), 6.79 (d, 1H, H22; J: 2,2 Hz), 6.85 (dd, 2H, H24, H26; J: 8.1, 2.2 Hz), 7.30 (t, 1H, H25; 8.1 Hz), 7.43 (dd, 1H, H15; J: 8.4, 1.9 Hz), 7.55 (d, 1H, H13; J: 1.9 Hz), 8.13 (d, 1H, H16; J: 8.4 Hz), 9.97 (brs, 1H, NH).

The elemental analysis was performed on a Perkin Elmer 2400 CHN analyzer; results (%) were within  $\pm 0.4$  of predicted values. Calculated for  $\text{C}_{17}\text{H}_{14}\text{N}_2\text{O}_3\text{S}_2\text{Cl}_2$ : C, 47.55 H, 3.29 N, 6.52. Found: C, 47.55; H, 3.29; N, 6.52.

The melting point (uncorrected) was measured with a Fischer-Johns micro hot-stage apparatus: 136–138 °C.

## Refinement

Hydrogen atoms were placed in calculated positions using a riding atom model with fixed C—H distances [0.86 Å for N, 0.93 Å for  $\text{C}(\text{sp}^2)$ , 0.96 Å for  $\text{C}(\text{sp}^3)$ ] and  $U_{\text{iso}} = p U_{\text{eq}}(\text{parent atom})$  [ $p = 1.2$  for N and  $\text{C}(\text{sp}^2)$ , 1.5 for  $\text{C}(\text{sp}^3)$ ].

## Figures

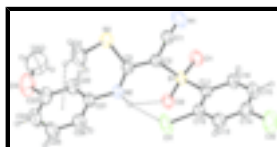


Fig. 1. Molecular structure of (I) showing the atomic numbering. Displacement ellipsoids are drawn at 50% probability level. Possible hydrogen bonds are shown as dashed lines.

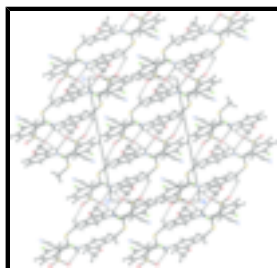


Fig. 2. Crystal structure of (I) viewed down the  $b$  axis. Possible hydrogen bonds are shown as dashed lines.

(2E)-2-(2,4-Dichlorophenylsulfonyl)-3-(3-methoxyanilino)-3-(methylsulfonyl)acrylonitrile

*Crystal data*

$C_{17}H_{14}Cl_2N_2O_3S_2$

$M_r = 429.32$

Monoclinic,  $P2_1/n$

Hall symbol: -P 2yn

$a = 11.6125 (7) \text{ \AA}$

$b = 10.1419 (6) \text{ \AA}$

$c = 16.4048 (10) \text{ \AA}$

$\beta = 100.926 (1)^\circ$

$V = 1897.0 (2) \text{ \AA}^3$

$Z = 4$

$F_{000} = 880$

$D_x = 1.503 \text{ Mg m}^{-3}$

Melting point = 409–411 K

Mo  $K\alpha$  radiation

$\lambda = 0.71073 \text{ \AA}$

Cell parameters from 3492 reflections

$\theta = 2.4\text{--}28.1^\circ$

$\mu = 0.58 \text{ mm}^{-1}$

$T = 296 (2) \text{ K}$

Prism, colourless

$0.29 \times 0.22 \times 0.17 \text{ mm}$

*Data collection*

Bruker SMART APEX  
diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

Detector resolution:  $8.13 \text{ pixels mm}^{-1}$

$T = 296(2) \text{ K}$

$\varphi$  and  $\omega$  scans

Absorption correction: multi-scan  
(SADABS; Bruker, 2001)

$T_{\min} = 0.796$ ,  $T_{\max} = 0.906$

12828 measured reflections

4659 independent reflections

3653 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.024$

$\theta_{\max} = 29.1^\circ$

$\theta_{\min} = 2.0^\circ$

$h = -13 \rightarrow 15$

$k = -13 \rightarrow 13$

$l = -22 \rightarrow 15$

*Refinement*

Refinement on  $F^2$

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.044$

$wR(F^2) = 0.119$

$S = 1.04$

4659 reflections

237 parameters

Primary atom site location: structure-invariant direct  
methods

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring  
sites

H-atom parameters constrained

$$w = 1/[\sigma^2(F_o^2) + (0.0578P)^2 + 0.5461P]$$

where  $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\max} = 0.001$

$\Delta\rho_{\max} = 0.58 \text{ e \AA}^{-3}$

$\Delta\rho_{\min} = -0.27 \text{ e \AA}^{-3}$

Extinction correction: none

# supplementary materials

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## Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

## Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
S1	0.77042 (4)	0.46268 (5)	0.02088 (3)	0.03340 (13)
S2	0.90726 (5)	0.19039 (6)	0.21790 (4)	0.05277 (17)
Cl1	0.95425 (5)	0.66027 (6)	0.13868 (5)	0.0661 (2)
Cl2	0.62331 (7)	1.02603 (7)	0.11885 (6)	0.0852 (3)
O1	0.66548 (12)	0.40951 (15)	-0.02832 (9)	0.0479 (4)
O2	0.87384 (12)	0.46887 (14)	-0.01526 (9)	0.0422 (3)
O3	1.2713 (2)	-0.0272 (2)	0.10486 (15)	0.0869 (7)
N1	0.62151 (18)	0.3565 (3)	0.17764 (15)	0.0692 (6)
N2	1.01127 (14)	0.34675 (17)	0.11897 (11)	0.0432 (4)
H2	1.0073	0.4157	0.0880	0.052*
C11	0.73323 (15)	0.62359 (19)	0.04938 (11)	0.0337 (4)
C12	0.81194 (16)	0.7078 (2)	0.09891 (13)	0.0395 (4)
C13	0.77859 (19)	0.8317 (2)	0.11955 (14)	0.0462 (5)
H13	0.8314	0.8872	0.1529	0.055*
C14	0.6645 (2)	0.8717 (2)	0.08946 (15)	0.0500 (5)
C15	0.58494 (19)	0.7912 (2)	0.04044 (15)	0.0524 (6)
H15	0.5086	0.8200	0.0214	0.063*
C16	0.61911 (17)	0.6676 (2)	0.01975 (13)	0.0425 (5)
H16	0.5659	0.6132	-0.0141	0.051*
C21	1.12266 (16)	0.2827 (2)	0.13590 (12)	0.0381 (4)
C22	1.13568 (19)	0.1547 (2)	0.11007 (13)	0.0462 (5)
H22	1.0715	0.1074	0.0822	0.055*
C23	1.2473 (2)	0.0979 (2)	0.12682 (15)	0.0534 (6)
C24	1.3416 (2)	0.1696 (3)	0.16728 (16)	0.0594 (7)
H24	1.4156	0.1311	0.1787	0.071*
C25	1.3275 (2)	0.2962 (3)	0.19075 (16)	0.0573 (6)
H25	1.3921	0.3441	0.2174	0.069*
C26	1.21789 (19)	0.3539 (2)	0.17530 (14)	0.0479 (5)
H26	1.2085	0.4406	0.1914	0.058*
C1	0.70534 (17)	0.3590 (2)	0.15151 (13)	0.0433 (5)
C2	0.80432 (16)	0.37253 (19)	0.11286 (12)	0.0352 (4)
C3	0.91295 (16)	0.31489 (19)	0.14429 (12)	0.0351 (4)
C4	1.0420 (2)	0.1996 (3)	0.29231 (15)	0.0709 (8)

H4B	1.1041	0.1607	0.2692	0.106*
H4C	1.0336	0.1527	0.3417	0.106*
H4A	1.0604	0.2902	0.3058	0.106*
C5	1.1784 (4)	-0.1029 (4)	0.0621 (3)	0.1160 (15)
H5B	1.1195	-0.1123	0.0956	0.174*
H5C	1.2069	-0.1884	0.0506	0.174*
H5A	1.1452	-0.0599	0.0109	0.174*

Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
S1	0.0278 (2)	0.0382 (3)	0.0323 (2)	0.00214 (17)	0.00072 (17)	-0.00024 (18)
S2	0.0432 (3)	0.0603 (4)	0.0538 (3)	0.0042 (2)	0.0066 (2)	0.0228 (3)
C11	0.0329 (3)	0.0604 (4)	0.0935 (5)	0.0095 (2)	-0.0169 (3)	-0.0257 (3)
C12	0.0768 (5)	0.0594 (4)	0.1172 (7)	0.0303 (3)	0.0126 (4)	-0.0190 (4)
O1	0.0398 (8)	0.0511 (9)	0.0462 (8)	-0.0016 (7)	-0.0088 (6)	-0.0057 (7)
O2	0.0385 (7)	0.0509 (8)	0.0386 (8)	0.0085 (6)	0.0112 (6)	0.0056 (6)
O3	0.0812 (14)	0.0708 (13)	0.1052 (17)	0.0331 (11)	0.0093 (12)	-0.0160 (12)
N1	0.0398 (11)	0.1008 (18)	0.0705 (15)	0.0078 (11)	0.0196 (10)	0.0213 (13)
N2	0.0305 (8)	0.0460 (9)	0.0531 (10)	0.0094 (7)	0.0075 (7)	0.0172 (8)
C11	0.0270 (8)	0.0384 (10)	0.0349 (10)	0.0035 (7)	0.0041 (7)	0.0024 (7)
C12	0.0291 (9)	0.0442 (11)	0.0435 (11)	0.0048 (8)	0.0024 (8)	-0.0011 (8)
C13	0.0438 (11)	0.0434 (12)	0.0505 (13)	0.0021 (9)	0.0063 (9)	-0.0060 (9)
C14	0.0502 (13)	0.0426 (12)	0.0590 (14)	0.0156 (10)	0.0149 (11)	0.0019 (10)
C15	0.0348 (11)	0.0563 (14)	0.0642 (15)	0.0155 (10)	0.0045 (10)	0.0072 (11)
C16	0.0294 (10)	0.0494 (12)	0.0459 (12)	0.0042 (8)	0.0002 (8)	0.0047 (9)
C21	0.0315 (9)	0.0460 (11)	0.0371 (10)	0.0083 (8)	0.0070 (8)	0.0083 (8)
C22	0.0419 (11)	0.0503 (12)	0.0445 (12)	0.0029 (9)	0.0036 (9)	0.0009 (9)
C23	0.0580 (14)	0.0526 (13)	0.0506 (13)	0.0206 (11)	0.0126 (11)	0.0014 (10)
C24	0.0366 (12)	0.0815 (19)	0.0593 (15)	0.0201 (11)	0.0074 (10)	0.0067 (13)
C25	0.0351 (11)	0.0728 (16)	0.0597 (15)	0.0007 (11)	-0.0020 (10)	0.0049 (12)
C26	0.0393 (11)	0.0524 (13)	0.0502 (13)	0.0033 (9)	0.0038 (9)	0.0023 (10)
C1	0.0321 (10)	0.0517 (12)	0.0447 (12)	0.0028 (9)	0.0040 (8)	0.0096 (9)
C2	0.0285 (9)	0.0408 (10)	0.0358 (10)	0.0011 (7)	0.0045 (7)	0.0041 (8)
C3	0.0309 (9)	0.0383 (10)	0.0347 (10)	0.0012 (7)	0.0031 (7)	0.0023 (8)
C4	0.0549 (15)	0.113 (2)	0.0428 (13)	0.0180 (15)	0.0042 (11)	0.0260 (14)
C5	0.118 (3)	0.073 (2)	0.150 (4)	0.002 (2)	0.009 (3)	-0.040 (2)

Geometric parameters ( $\text{\AA}$ ,  $^\circ$ )

S1—O1	1.4333 (14)	C15—H15	0.9300
S1—O2	1.4380 (14)	C16—H16	0.9300
S1—C2	1.7443 (19)	C21—C26	1.375 (3)
S1—C11	1.7733 (19)	C21—C22	1.383 (3)
S2—C3	1.757 (2)	C22—C23	1.397 (3)
S2—C4	1.795 (3)	C22—H22	0.9300
C11—C12	1.7264 (19)	C23—C24	1.375 (4)
C12—C14	1.732 (2)	C24—C25	1.359 (4)
O3—C23	1.362 (3)	C24—H24	0.9300



## supplementary materials

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O3—C5	1.399 (4)	C25—C26	1.380 (3)
N1—C1	1.136 (3)	C25—H25	0.9300
N2—C3	1.327 (2)	C26—H26	0.9300
N2—C21	1.427 (2)	C1—C2	1.421 (3)
N2—H2	0.8600	C2—C3	1.397 (3)
C11—C12	1.394 (3)	C4—H4B	0.9600
C11—C16	1.396 (2)	C4—H4C	0.9600
C12—C13	1.376 (3)	C4—H4A	0.9600
C13—C14	1.384 (3)	C5—H5B	0.9600
C13—H13	0.9300	C5—H5C	0.9600
C14—C15	1.373 (3)	C5—H5A	0.9600
C15—C16	1.377 (3)		
O1—S1—O2	118.14 (9)	C21—C22—H22	120.7
O1—S1—C2	108.56 (9)	C23—C22—H22	120.7
O2—S1—C2	108.08 (9)	O3—C23—C24	115.9 (2)
O1—S1—C11	105.74 (9)	O3—C23—C22	124.1 (2)
O2—S1—C11	109.43 (9)	C24—C23—C22	120.1 (2)
C2—S1—C11	106.30 (9)	C25—C24—C23	120.6 (2)
C3—S2—C4	106.78 (12)	C25—C24—H24	119.7
C23—O3—C5	117.8 (2)	C23—C24—H24	119.7
C3—N2—C21	129.07 (17)	C24—C25—C26	120.3 (2)
C3—N2—H2	115.5	C24—C25—H25	119.8
C21—N2—H2	115.5	C26—C25—H25	119.8
C12—C11—C16	118.77 (18)	C21—C26—C25	119.7 (2)
C12—C11—S1	123.49 (14)	C21—C26—H26	120.2
C16—C11—S1	117.73 (15)	C25—C26—H26	120.2
C13—C12—C11	121.26 (18)	N1—C1—C2	173.9 (2)
C13—C12—Cl1	117.11 (16)	C3—C2—C1	123.04 (18)
C11—C12—Cl1	121.62 (15)	C3—C2—S1	125.34 (15)
C12—C13—C14	118.3 (2)	C1—C2—S1	111.57 (14)
C12—C13—H13	120.8	N2—C3—C2	123.57 (17)
C14—C13—H13	120.8	N2—C3—S2	122.47 (14)
C15—C14—C13	121.8 (2)	C2—C3—S2	113.91 (14)
C15—C14—Cl2	120.62 (17)	S2—C4—H4B	109.5
C13—C14—Cl2	117.50 (18)	S2—C4—H4C	109.5
C14—C15—C16	119.48 (19)	H4B—C4—H4C	109.5
C14—C15—H15	120.3	S2—C4—H4A	109.5
C16—C15—H15	120.3	H4B—C4—H4A	109.5
C15—C16—C11	120.3 (2)	H4C—C4—H4A	109.5
C15—C16—H16	119.8	O3—C5—H5B	109.5
C11—C16—H16	119.8	O3—C5—H5C	109.5
C26—C21—C22	120.81 (19)	H5B—C5—H5C	109.5
C26—C21—N2	118.09 (19)	O3—C5—H5A	109.5
C22—C21—N2	121.03 (18)	H5B—C5—H5A	109.5
C21—C22—C23	118.5 (2)	H5C—C5—H5A	109.5
O1—S1—C11—C12	179.07 (17)	C5—O3—C23—C22	0.7 (4)
O2—S1—C11—C12	-52.70 (19)	C21—C22—C23—O3	179.8 (2)
C2—S1—C11—C12	63.80 (19)	C21—C22—C23—C24	-0.8 (3)

O1—S1—C11—C16	-1.47 (18)	O3—C23—C24—C25	179.0 (2)
O2—S1—C11—C16	126.77 (16)	C22—C23—C24—C25	-0.5 (4)
C2—S1—C11—C16	-116.74 (16)	C23—C24—C25—C26	0.9 (4)
C16—C11—C12—C13	0.8 (3)	C22—C21—C26—C25	-1.3 (3)
S1—C11—C12—C13	-179.76 (17)	N2—C21—C26—C25	-178.4 (2)
C16—C11—C12—C11	179.47 (16)	C24—C25—C26—C21	0.0 (4)
S1—C11—C12—C11	-1.1 (3)	O1—S1—C2—C3	128.18 (18)
C11—C12—C13—C14	-0.4 (3)	O2—S1—C2—C3	-1.1 (2)
C11—C12—C13—C14	-179.12 (18)	C11—S1—C2—C3	-118.48 (18)
C12—C13—C14—C15	0.3 (4)	O1—S1—C2—C1	-49.26 (17)
C12—C13—C14—C12	178.07 (18)	O2—S1—C2—C1	-178.53 (15)
C13—C14—C15—C16	-0.6 (4)	C11—S1—C2—C1	64.08 (17)
C12—C14—C15—C16	-178.30 (18)	C21—N2—C3—C2	-169.2 (2)
C14—C15—C16—C11	1.0 (3)	C21—N2—C3—S2	8.0 (3)
C12—C11—C16—C15	-1.1 (3)	C1—C2—C3—N2	-166.4 (2)
S1—C11—C16—C15	179.43 (17)	S1—C2—C3—N2	16.5 (3)
C3—N2—C21—C26	-120.4 (2)	C1—C2—C3—S2	16.2 (3)
C3—N2—C21—C22	62.5 (3)	S1—C2—C3—S2	-160.99 (12)
C26—C21—C22—C23	1.7 (3)	C4—S2—C3—N2	37.0 (2)
N2—C21—C22—C23	178.65 (19)	C4—S2—C3—C2	-145.56 (17)
C5—O3—C23—C24	-178.8 (3)		

Hydrogen-bond geometry ( $\text{\AA}$ ,  $^\circ$ )

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
N2—H2 $\cdots$ O2	0.86	2.14	2.756 (2)	129
N2—H2 $\cdots$ O2 <sup>i</sup>	0.86	2.31	3.005 (2)	138
N2—H2 $\cdots$ C11	0.86	2.72	3.2763 (18)	123
C4—H4C $\cdots$ O1 <sup>ii</sup>	0.96	2.46	3.215 (3)	135
C16—H16 $\cdots$ O1	0.93	2.40	2.815 (3)	107
C4—H4B $\cdots$ Cg2	0.96	2.74	3.501 (3)	137

Symmetry codes: (i)  $-x+2, -y+1, -z$ ; (ii)  $x+1/2, -y+1/2, z+1/2$ .

Fig. 1

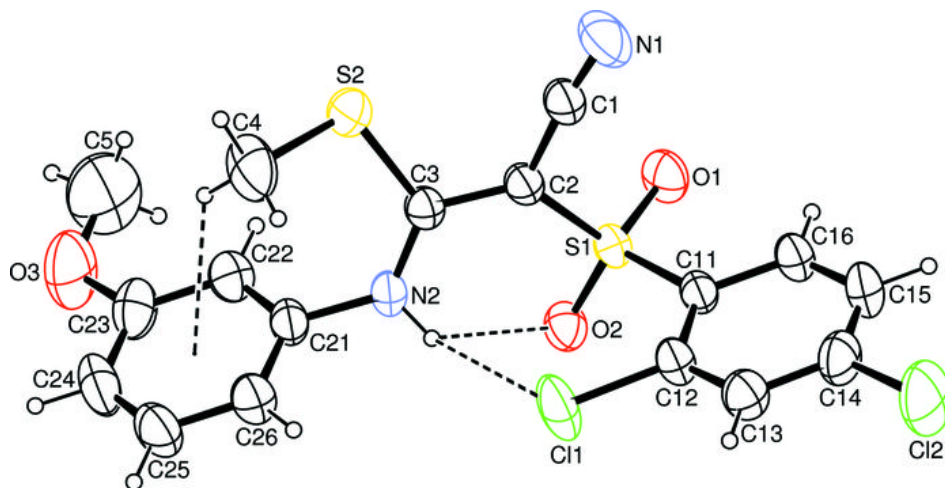


Fig. 2

